

MICROPARTICLES AND METHODS FOR THEIR PRODUCTION

FIELD OF THE INVENTION

5 This invention relates to microparticles having metal-containing cores encapsulated in coatings in the form of shells of graphite or silica, and to methods of making such microparticles.

BACKGROUND TO THE INVENTION

10 There has been great interest in the incorporation of foreign materials into enclosed nano-carbon cages. This interest has been driven by the potential applications of these filled carbon capsules, which lie in areas as diverse as optical, electronic, storage, magnetic recording materials, and nuclear medicine. In particular, carbon (onion-shell) structures of extreme strength may offer excellent protection to their encapsulated nanomaterials for applications. In addition, the close packing structures of the carbon shells do not allow exchange of molecules/materials from inner cores to the external environment. We have now found that materials of extremely reactive (air sensitive) or hazardous (radioactive) nature can be safely caged in the carbon-enclosed structures.

25 The idea of using carbon caged structures as molecular containers dates back many years [1,2] before the actual discovery of fullerene and related structures. Research in this area is chiefly driven by the potential applications of filled nanocapsules in areas such as electronics (quantum dots and wires), magnetic data storage, xerographics, lubrication, sensors and medicinal materials [3-5]. So far, several groups have succeeded

30

in encapsulating various nanosized materials into fullerene structures by either using Krätschmer-Huffman arc [6,7], laser [8] or electron irradiation methods [9] during the fullerene synthesis or by creating an
5 'opening' in the carbon caged structures using chemical methods [10-12] prior to filling. Yields of the endohedral metallofullerenes synthesis [8] tend to be low and there are doubts about stability. Filled polygonal shaped carbon nanoparticles and nanotubes appear to be
10 more promising from the point of view of applications. The techniques currently used to synthesise these structures include modifications of Krätschmer-Huffman arc method and methods involving heat treatment of impregnated microporous carbon [13,14]. It is noted that
15 there is no macroscopic synthetic method yet developed for the preparation of filled carbon onions (quasi-spherical graphitic shells) although minuscule amounts were produced through irradiation of amorphous carbon, nanotubes or nanoparticles with foreign atoms with an
20 intense electron beam [9]. No reported work has been found on the preparation of total carbon encapsulation of radioactive nanoparticles.

25 SUMMARY OF THE INVENTION

By this invention, a novel methodology for the encapsulation of radioactive elements within the carbon cage is provided. This simple technique allows production of macroscopic quantities of quasi-spherical
30 graphitic/fullerenic shells (carbon onions) that can encapsulate nanoparticles containing radioactive element(s) of a very narrow range of particle diameters.

We believe the method may offer new routes for safe handling/disposal of radioactive substances and the carbon coated nano-radioactive products may find applications in medical imaging or radiotherapy.

- 5 According to the invention in a first aspect, there is provided a method of making microparticles having a metal-containing core encapsulated in a graphitic shell containing hetero atoms, including the steps of
- 10 (i) forming, in a liquid medium, colloidal particles containing a first metal in the form of a metal-oxo species, the first metal being selected from Fe, Co, Ni and Pd, the particles being colloidally stabilized by a surfactant and containing in addition to the first metal source material of carbon and the hetero atoms, and
- 15 (ii) separating said particles from the liquid medium and pyrolyzing them in inert gas to yield the microparticles having said core and said graphitic shell containing said hetero atoms encapsulating the core.

20 The basis of the invention in this aspect are the findings that stable colloidal particles can be produced containing metal-oxo species of Fe, Co, Ni or Pd, and that these particles can be pyrolyzed to provide a metal-containing core, typically of metal, alloy, carbide or oxide or mixtures of these, encapsulated in a good

25 quality graphitic shell which contains hetero atoms which are believed to cause curvature to the carbon-atom layers of the graphite. Further it has been found that a second metal, such as a radionuclide, can be incorporated in the colloidal particles. The first metal, which is at least

30 one of Fe, Co, Ni and Pd, is believed to play a role in catalyzing the formation of the graphite during pyrolysis. The hetero atom, which should be chosen to

become chemically bound in the graphite molecular structure, may be selected from N, B, P, S and O, and is most preferably N. The hetero atom is believed to cause curvature of the carbon layers of the graphitic shell, resulting in an encapsulating coating having few or no defects or fractures.

The metal-oxo species in the colloidal particle is typically insoluble in the medium which contains it. It may be produced by oxidation of a soluble compound of the first metal. Two methods of carrying this out are proposed.

In one particular method, step (i) is performed by forming a solution of a compound of the first metal in the liquid medium which is a polar solvent and heating the solution in the presence of the surfactant and an oxidizing agent to convert the compound to the metal-oxo species and form the colloidal particles. In this method, a high boiling point polar solvent is preferably employed, e.g. one having a boiling point above 100°C, or even above 200°C. The oxidizing agent may be for example oxygen or a compound of a second metal which becomes incorporated in the colloidal particles. Such a compound of a second metal is for example an oxo-anion, such as ReO_4^- or TcO_4^- .

In a second method, step (i) is performed by forming an emulsion having dispersed phase droplets stabilized by the surfactant and containing a dissolved compound of the first metal and oxidizing the compound of the first metal to produce the metal-oxo species and form the colloidal particles. The dispersed phase droplets may include a dissolved oxidizing agent in the form of a compound of a second metal, as described above.

The colloidal particle includes source material of carbon and the hetero atom, for the formation of the graphitic shell. Thus it is not required that another carbon source, e.g. a gas, is provided during the
5 pyrolysis. The carbon may effect reduction of the metal species, to provide metal or alloy as the core. The carbon and hetero atom may be present in a complex of the first metal. Preferably this metal complex includes ligands selected from cyanide, isocyanide, cyanate and
10 isocyanate.

In a second aspect, the invention provides microparticles having a core containing a radionuclide encapsulated by a graphitic shell. When produced by the methods described above, the core contains in addition to
15 the radionuclide, at least one of Fe, Co, Ni and Pd. Further the graphitic shell may contain chemically bound hetero atoms such that the graphitic carbon layers of the shell are curved.

In nuclear medicine, the choice and form of
20 radionuclides should be carefully chosen. The choice of a radionuclide for imaging purposes is chiefly dictated by the necessity of minimizing the radiation dose to the patient and the detection characteristics of present-day nuclear medicine instrumentation. The forms should
25 preferably be non-toxic in the desired amounts, and would not directly involve in the patient's physiological body mechanisms. The form, structure and morphology and concentration of the radioactive material will also affect the imaging quality (small particle size will give
30 better imaging quality i.e. smaller pixels). This invention describes a novel, easy, convenient method for the synthesis of complete carbon encapsulated

radionuclides with a narrow particle size distribution
suitable for diagnostic imaging and therapeutic
applications. The impermeable carbon coatings offer
total isolation of the enclosed nuclides from the
5 environment.

The novel method described herein is concerned with
a methodology for the production of a complete carbon
encapsulated radioactive material with a small range of
particle size. We believe these materials produced from
10 the method will find applications in lung ventilation and
lung perfusion imaging and other diagnostic and
radiotherapy areas.

This method may be suitable to encapsulate a wide
range of radionuclides (parental or daughter nuclides)
15 with different emitted radiation and decay times.

20	Typical example, (i)	^{99}Mo	\rightarrow	$^{99\text{m}}\text{Tc}$	\rightarrow	^{99}Tc	\rightarrow	^{99}Ru
	Half-life:	67 hours		6 hours		long		stable
25	(ii)	^{113}Sn	\rightarrow	$^{113\text{m}}\text{Sn}$	\rightarrow	^{113}In		
	Half-life	115 days		1.67 hours		stable		
30	(iii)	^{68}Ge	\rightarrow	^{68}Ga	\rightarrow	^{68}Zn		
	Half-life	275 days		1.1 hours		stable		
35	(iv)	^{81}Rb	\rightarrow	$^{81\text{m}}\text{Kr}$	\rightarrow	^{81}Kr		
	Half-life	4.7 hours		13 seconds		stable		

Technetium-99m ($^{99\text{m}}\text{Tc}$) has excellent physical
30 characteristics for diagnostic imaging having a half-life
of 6 h and emitting gamma-ray photons at an energy of 140
keV which are suitable for detection with a gamma camera.
 $^{99\text{m}}\text{Tc}$ is easily obtained from a commercially available
 $^{99}\text{Mo}/^{99\text{m}}\text{Tc}$ generator. ^{99}Mo is a fission product which is
35 obtainable in a carrier-free form with a high specific

activity. It has a half-life of 66.02 h and decays to ^{99m}Tc by β^- emission, which decays by isomeric transition to ^{99}Tc . In commercially available generators ^{99}Mo is absorbed onto an alumina column. Elution of the generator with 0.9% sodium chloride solution elutes the ^{99m}Tc in the form of sodium pertechnetate ($\text{Na}^{99m}\text{TcO}_4$) leaving the ^{99}Mo bound to the column¹⁵. $^{99}\text{Mo}/^{99m}\text{Tc}$ generators are generally purchased by hospitals on a weekly basis and have a working life of one week. ^{99m}Tc is incorporated in a variety of chemical forms for oral intake, injection or for lung ventilation/perfusion.

For lung imaging "Technegas" has been developed. It is an ultrafine ^{99m}Tc -labelled aerosol introduced by W. Burch [22] in 1986. The small radioactive particles (nanometric in size) in a Technegas aerosol result in better clinical images than those obtained from other radioactive aerosols. It is known that the size and the content of the Tc particle in the aerosol can critically affect the quality of the images. The size of the aerosol particles also determines the site of lung deposition (bronchial and alveolar regions), subsequent modes and rates of clearance and hence will affect the diagnostic information obtained. More recent work, using electron microscopy, has shown that ^{99m}Tc particles in Technegas are in fact relatively large crystals of over 100 nm in diameter. Another problem that arises with the production of Technegas is that significant quantities of ^{99m}Tc are incompletely protected in the carbon matrix. This leads to leaching of radioactive ^{99m}Tc to body fluids hence resulting in the undesirable appearance of ^{99m}Tc in the saliva, oesophagus, stomach and thyroid gland and thus degrading the quality of the images. In addition

the high cost of purchasing the commercial Technegas generator has limited its use in many centres.

Our method in one embodiment involves the preparation of graphitic-encapsulated microparticles, more particularly nanoparticles, by the dissolution of a source or sources of iron, carbon and a hetero atom (usually N, but B, P, S or O for example may be used) in a polar solvent at elevated temperature; partially decomposing the iron compound to an insoluble iron oxygen species; size controlling and stabilizing the partially oxidised iron species by use of a surfactant; and pyrolyzing the stabilized species to yield a metallic core encapsulated by a graphitic/hetero shell. To form a radionuclide-containing particle, a salt or complex of a radionuclide is also present. It is found that during the partial decomposition of the iron compound, especially in the absence of oxygen, the radionuclide is incorporated into the core.

Since the graphitic encapsulation derives primarily from the carbon source, regulation of the carbon:iron ratio can help determine the numbers of encapsulating layers around a core; this in turn may act as a modulator or regulator of the radiation emitted from a radionuclide in the core.

In more detail some of the most important synthesis steps in obtaining the desired products have been elucidated from study of an iron cyanide complex as a starting material. These include the fact that iron species (catalyst) catalyse formation of enclosed graphitic structure from its attached cyanide ligand (or carbon and nitrogen containing ligand, stabiliser) upon heating, which will subsequently enclose the iron

particle forming completely encapsulated particles.

Prior to heat treatment, control in particle size is important in order to make encapsulated nanosized particles. We therefore show that adding surfactants or
5 polymers controls the particle sizes and stabilisation of insoluble iron and cyanide-containing nanoparticles when soluble iron cyanide species is partially decomposed in a high-boiling polar solvent. This is based on the fact that the polymer/surfactant adsorbs at the surface of the
10 newly formed nano-particles forming micellar protection so preventing aggregation. In other contexts, stabilisation of colloid in a polymer solution is a well-established technique typically in paint and inkjet technologies. Thus, we show that controlled size of Fe
15 cyanide containing nanoparticles stabilised with adsorbed polymers/surfactants is a most important step before the radioactive element incorporation. The incorporation of foreign element(s) to the stabilised iron containing particles can be subsequently performed by developing
20 specific interactions (chemical linkage, ion-exchange, redox attachment, etc) with the Fe(II) containing particles. Finally, after the removal of solvents followed by heat treatment the stabilised nanoparticles containing iron and foreign atoms with a narrow size
25 distribution will give rise to the desired products. It is noted that the residue cyanide species in the particles will provide the carbon source for the formation of enclosed graphitic structure. We show that Re and Tc species can be successfully incorporated into
30 the carbon-encapsulated iron containing nanoparticles via using ReO_4^- or TcO_4^- . Incorporation of $^{99\text{m}}\text{Tc}$ is thus possible. Re and Tc have almost identical chemical

properties (due to the lanthanide contraction). Other radioactive elements such as ^{99m}Mo , ^{113}Sn can similarly be incorporated into this novel carrier (surfactant stabilised iron(oxo) cyanide nanoparticles) by employing
5 similar chemistry.

The invention thus provides a low temperature solution method for encapsulation of radioactive nanoparticles/clusters with a fine control in particle size. There are many advantages in the safe handling/disposal
10 of radioactive materials if the materials are stored in an impervious carbon coating. Early work suggested it was difficult to carry out complete encapsulation of material with graphitic carbon unless the temperature rose above the vaporisation temperatures of graphite
15 ($>2500^\circ\text{C}$) during electric arc excitation [22]. However, there are many disadvantages associated with the extremely high temperature synthetic method especially regarding to the difficulty in controlling the particle size. Complete carbon encapsulation over radioactive
20 nanoparticles thus may require a high temperature since amorphous carbon atoms do not offer total protection to these particle against leaching. The crystallisation of amorphous carbon atoms into complete graphitic protective capsule structures (graphitic carbon shells) is known to
25 take place at minimum temperature of 2500°C [25]. However, the alternative would be to induce graphitisation at low temperatures by the use of catalyst. The phenomenon of low temperature ($<700^\circ\text{C}$) "catalytic graphitisation" from a different variety of
30 carbon sources over Fe, Co, Ni, Cr, Pd, has been known for many years. Here we demonstrate clearly that the formation of iron or iron carbide encapsulated in carbon

shells from heat treatments of stabilised iron (oxo) cyanide species can be achieved. The cyanide species provide carbon source for the formation of enclosed graphitic shells. This indicates an effective catalytic carbonisation over the iron materials.

In summary, we report the (a) effectiveness of using catalytic component (Fe or alternatively Ni, Co or Pd, e.g. as cyanides) to form well-defined sized nanoparticles in a high boiling polar solvent (with stabiliser) during their cyanide decomposition; (b) incorporation of foreign atoms (radioactive atom, such as ^{99m}Tc) into the iron oxo-cyanide nanoparticle aggregates can be achieved using the oxidative $^{99m}\text{TcO}_4$ ions through a fast redox trapping mechanism; (c) thermal decomposition of these colloidal particles (with cyanide) result in the enclosed graphitic structure embracing radioactive element(s).

In another aspect, the invention provides the formation of silica-coated microparticles, by a similar method. In this aspect, the invention provides a method of making solid microparticles having a metal-containing core surrounded by a silica coating, including the steps of

- (i) forming, in a liquid medium, colloidal particles containing a metal-containing species and colloiddally stabilized by a surfactant, and
- (ii) forming a silica coating around said colloidal particles by hydrolyzing a silicon compound in the region of the interface between the colloidal particle and the liquid medium.

A plurality of metal-containing species of different metals may be included in the colloidal particles, and

thus in the core of the microparticles produced.

Typically said metal-containing species is selected from metal, alloy, metal oxide, metal hydroxide and metal carbide. Preferably the metal-containing species is
5 ferromagnetic (enabling magnetic separation of the microparticles from liquid) and/or contains a radionuclide.

Preferably the microparticles are aged, e.g. for a day or more, before removal from the system containing
10 the silicon compound, in order to establish the silica coating to the desired thickness. The coating may be porous.

In a preferred form of the method in step (i) the colloidal particles are made by forming an emulsion
15 having dispersed phase droplets stabilized by the surfactant and containing a dissolved compound of the metal and causing the metal-containing species to precipitate thereby forming the colloidal particles. The precipitation may be caused by addition of alkali.

20 The silicon compound which is hydrolyzed may be an alkoxy silane compound, i.e. a compound containing at least one Si-OR linkage, where R is alkyl of preferably 1 - 8 carbon atoms, more preferably 1 - 4 carbon atoms, such as tetraethyl ortho silane (TEOS, $\text{Si}(\text{OC}_2\text{H}_5)_4$).

25 Further the invention provides microparticles each having a core comprising at least one metal-containing species which is ferromagnetic and/or contains a radionuclide and a coating of silica gel encapsulating the core, which may be produced by the above method. The
30 silica gel may have at its surface functional groups, e.g. OH groups, for the attachment of other species, such

as biochemical or biological species (e.g. peptides, markers, cognate binding partner, solubilizers).

The cores of the silica-coated microparticles preferably have an average diameter in the range 1 to 100 nm, more preferably 1 to 50 nm. The silica gel coating may have any desired thickness, but preferably has an average thickness in the range 1 to 50 nm, preferably 2 to 10 nm, e.g. 2 to 4 nm.

By this method, ferromagnetic cores and/or radionuclides can be encapsulated in silica coatings.

BRIEF INTRODUCTION OF THE DRAWINGS

In the drawings:

Figure 1 is an electron micrograph showing a large discrete iron-containing particle encapsulated within multi-layer graphitic carbon shell prepared from decomposition of iron cyanide particle in nitrogen (scale bar = 50 nm);

Figure 2 shows the powder X-ray diffraction (XRD) spectrum of pure $\text{Fe}^{\text{III}}_4[\text{Fe}^{\text{II}}(\text{CN})_6]_3$ salt acting as a precursor;

Figures 3 and 4 show the powder X-ray diffraction (XRD) spectra of partial air decomposed $\text{Fe}^{\text{III}}_4[\text{Fe}^{\text{II}}(\text{CN})_6]_3$ in refluxing dioctylether (Figure 3: small amount of dissolved oxygen in dioctylether; Figure 4: oxygen continuously purging through the system);

Figure 5 (a TEM micrograph) shows the colloidal stable iron-oxo-cyanide containing nanoparticles obtained via refluxing 0.86 g $\text{Fe}^{\text{III}}_4[\text{Fe}^{\text{II}}(\text{CN})_6]_3$ salt in 40 ml dioctylether solvent with addition of oleic acid (20 ml). Small but homogeneous sized organic stabilised

nanoparticles (~ 10 nm) are clearly visible (nanosize particles giving a super-lattice packing);

Figure 6a is a typical high resolution transmission electron micrograph (lattice image) showing iron nanoparticles with the interplanar spacing of 0.21 ± 0.05 nm corresponding to (111) of fcc Fe (cementite) encapsulated in concentric quasi-spherical graphitic shells of 3.4×10^{-10} m spacing (scale bar = 5 nm);

Figure 6b is a TEM showing the same product as Figure 6a on a larger scale (scale bar = 1 nm);

Figure 7 is the EDX spectrum of Re-incorporated stabilised Fe nanoparticles which show the selected area rich in iron and rhenium (Cu peaks arise from the copper grids used for support);

Figure 8 shows the corresponding high-resolution TEM image of Re-incorporated stabilised Fe nanoparticles with very well defined size and shape (as Figure 7) (scale bar = 20 nm);

Figure 9a is a high resolution TEM image of a spherical graphitic shell structure (2-5 layers) filled with Fe and Re containing nanoparticles after applying heat treatments to the particles of Figure 8 (scale bar = 20 nm);

Figure 9b is a high resolution TEM showing the same product as Figure 9a on a larger scale (scale bar = 5 nm), with inset EDX spectrum of the selected area showing Re and Fe in the particle core (the Cu peaks arise from copper support grids);

Figure 10 shows the result of irradiation of amorphous carbon with intense electron beam in TEM showing it to have spherical graphitic concentric shells

(onions); and a diagrammatic model attached (scale bar 100×10^{-10} m);

Figure 11 is a high resolution TEM showing a spherical graphitic shell structure filled with a nanoparticle containing Fe and ^{99m}Tc produced in Example 4;

Figure 12 is a bar chart of activity counts obtained in the procedure described in Example 4;

Figure 13 is the XRD spectrum of the iron oxide core nanoparticles of Example 5;

Figure 14 is the XRD spectrum of the iron cobalt oxide core nanoparticles of Example 6;

Figure 15 is the EDS analysis for the iron oxide core nanoparticles of Example 5;

Figure 16 is the EDS analysis for the iron cobalt oxide core nanoparticles of Example 6; and

Figure 17 is a TEM micrograph of the particles of Example 5 (scale bar = 5 nm).

Experimental Details and Results

Example 1: Synthesis of Carbon Encapsulated Iron Particles (comparative)

A sample of iron (III) ferrocyanide (purity >99.9%) obtained from Sigma plc, was transferred into a quartz tube with one end plugged with quartz wool. The filled tube, which was placed in the tube furnace, was first heated to 200°C (heating rate: 5°C/min, duration: 2h), and then finally to 900°C (heating rate: 10°C/min; duration: 3h) in N_2 atmosphere (approx. flow rate: 0.9 l/min). The black materials formed were stored in a sample vial. Elemental analyses from the EDX analysis and microanalysis showed that the sample after the heat

treatment contained mainly C, Fe and a small amount of nitrogen (2-4%), which was scratched off the tube and examined by TEM. The TEM micrograph in Figure 1 shows that an iron-containing particle (iron and iron carbides) is encapsulated by graphitic carbon shells (onion structure) though the sizes of most of these encapsulated particles are very large ($>0.1\mu\text{m}$). Thus, it is clear that iron catalyses formation of enclosed graphitic structure from cyanide- or carbon-nitrogen-containing ligand. Other metals such as alkali or alkaline earths are unable to catalyse these enclosed graphitic structures from their cyanide salts, nor are iron species without using nitrogen containing salt/ligand/stabiliser.

Example 2: Synthesis of Carbon Encapsulated Iron Nanoparticles:

Example 1 discloses the use of iron cyanide salt for the formation of an enclosed graphitic structure embracing an iron particle, but no control in particle size was achieved. Here we describe the use of a polar high boiling solvent (such as dioctyl ether of boiling point 287°C) to dissolve the iron ferrocyanide compound at refluxing temperature (the compound is fairly soluble in the solvent at elevated temperature) giving intense blue colour solution. Iron-containing cyanide compounds are known to decompose at about $200\text{--}250^{\circ}\text{C}$ [23,24]. In the presence of dissolved oxygen (air) achieved by purging the system with an air stream, the $\text{Fe}^{\text{II}}(\text{CN})_6^{4-}$ of the $\text{Fe}^{\text{III}}_4[\text{Fe}^{\text{II}}(\text{CN})_6]_3$ salt decomposes to $\text{Fe}(\text{CN})_{2-3}\text{O}_{1-2}^{n-}$ losing the cyanide species but simultaneously replacing them with oxygen species [23]. This will ultimately lead to precipitating of iron oxide that is insoluble in the

solvent. Figure 2 shows the powder X-ray diffraction (XRD) of the pure $\text{Fe}^{\text{III}}_4[\text{Fe}^{\text{II}}(\text{CN})_6]_3$ salt. Iron oxide structure (magnetite, Fe_3O_4) is formed from the partial iron cyanide decomposition by air, giving mixed phases in the XRD patterns. The extent of the decomposition depends on the oxygen availability and duration of the treatment (Figures 3 and 4). An organic surfactant/stabiliser, such as oleic acid or polyvinylpyrrolidone (PVP), stabilises the partially oxidised iron cyanide species colloiddally against precipitation from the solution. Figure 5 (TEM micrograph) shows the colloidal stable iron-oxo-cyanide containing nanoparticles obtained via refluxing 0.86g $\text{Fe}^{\text{III}}_4[\text{Fe}^{\text{II}}(\text{CN})_6]_3$ salt in 40 ml dioctylether solvent with adding oleic acid (20 ml). Small but homogeneous particle size (nanosize particles giving a super-lattice packing) is obtained using this preparative mixture. There is a clear indication that the size of the particle can also be finely controlled by tailoring the ratio of oleic acid to the iron compound. Oleic acid, having a polar head group (acid group) and a double bond in the middle of the molecule, acts as surfactant. It is envisaged that the double bond will interact strongly with the iron species in the inner core of a micelle while its polar head will face outwards forming a normal phase micelle in this polar solvent. Hence the micellar structure provides a fine control in particle size. Application of the same heat treatment as mentioned in Example 1 to the colloidal iron containing particles of this example produces iron/iron carbide particle encapsulated in the enclosed graphitic carbon shelled

structure (see Figures 6a and 6b and compare the model view in Figure 6c).

Example 3: Synthesis of Carbon Encapsulated Nanoparticles containing Iron and Rhenium:

Example 2 shows that molecular oxygen from air can partially oxidise the iron cyanide compound in dioctylether at elevated temperatures. Adding a small quantity of sodium perrhenate (inorganic oxidant) to the mixture, as described in Table 1, can also oxidise the iron (II) cyanide species. Experimental details for typical synthesis are described below:

Table 1. The starting materials. All the materials were used as supplied by Aldrich.

Chemical Name	Chemical Formula	Mol.Wt.	Act. Wt/g	Features
Sodium Perrhenate	NaReO_4	273	0.28	White Crystals
Iron (III) Ferrocyanide	$\text{Fe}^{\text{III}}_4[\text{Fe}^{\text{II}}(\text{CN})_6]_3$	859.25	0.86	Blue powder
Dioctylether	$[\text{CH}_3(\text{CH}_2)_7]_2\text{O}$	242.45	40 ml	Clear Liquid
Oleic acid	$\text{CH}_3(\text{CH}_2)_7\text{CH}=\text{CH}(\text{CH}_2)_7\text{COOH}$	282	20 ml	-

The chemicals shown above, see Table 1, were mixed in a 3-necked round bottom flask using a magnetic stirrer bar. Before the mixture was heated up to 290°C for 18h, a gentle stream of nitrogen gas was bubbled directly into the mixture for 30 min. Nitrogen gas was continuously bubbled into mixture during the reaction. The black mixture formed was allowed to cool to room temperature. To separate the particles, 20 ml of ethanol was added to the product before the resultant mixture was centrifuged at 4000 rpm for 20 min. This procedure was repeated 3

times. The blue resultant solids were allowed to air dry in the fume cupboard.

Carbon encapsulation is achieved by heating the particles in nitrogen as follows. The as-synthesised sample was first transferred into a quartz tube with one end plugged with quartz wool. The filled tube, which was placed in the tube furnace, was first heated to 200°C (heating rate: 5°C/min, duration: 2h), and then finally to 900°C (heating rate: 10°C/min; duration: 3h) in N₂ atmosphere (approx. flow rate: 0.9 l/min). The black materials were stored in a sample vial.

After heat treatment (which may be carried out in various stages or steps) the closed graphitic structures are observed embracing both the iron and rhenium elements within the carbon enclosed shells (Figure 9).

Hence, during the decomposition of iron cyanide at elevated temperature, in the absence of air, the oxygen species will be transferred from the ReO_4^- species to the $\text{Fe}^{\text{II}}(\text{CN})_6^{4-}$ (Re(+7) will be reduced while simultaneously oxidising the Fe^{2+}) leading to incorporation of the Re species into the stabilised very well-defined, nano-sized iron (oxo) cyanide particles (see EDX analysis in Figure 7 and the TEM micrograph in Figure 8 prior heat treatments).

Because of the d- electron configuration (lanthanide contraction) rhenium (Re) shows almost identical chemical properties as technetium (Tc). Typically, the reduction potentials of the ReO_4^- and $^{99\text{m}}\text{TcO}_4^-$ species are about 0.7 V (V versus NHE). It is reported that a fast rate of reductive deposition of TcO_4^- is obtained over small solid magnetite (iron oxide). $^{99\text{m}}\text{TcO}_4^-$, through its reductive deposition onto the organic stabilised nano-iron

containing particles as described in this process, followed by heat treatment, will make a carbon encapsulated particle containing radioactive ^{99m}Tc element. As a result, the small but defined dimensions of the iron core (as a catalyst) can incorporate a significant amount of radioactive ^{99m}Tc into the final carbon encapsulated particles for imaging, storage and radio-therapeutic applications by this technique.

Example 4 : Synthesis of Carbon Encapsulated Nanoparticles containing Iron and Technetium:

The procedure of Example 3 was repeated, using $\text{Na}^{99}\text{TcO}_4$ in place of NaReO_4 . The technetium was incorporated into the colloidal particles formed and appeared in the pyrolyzed graphitic shell particles. Fig. 11 is a TEM image of the spherical graphitic shell nanoparticles filled with Fe and ^{99m}Tc in the final product. Fig. 12 shows activity counts at different stages of the procedure, as now described. Activity assays were evaluated using gamma counter: 1 mL of $^{99m}\text{TcO}_4^-$ of 5.19×10^{-11} mole dm^{-3} as the standard. This solution added to the mixture as described and allowed to reflux for 1 h (~1 mL water collected with no radioactivity). Ethanol addition lead to precipitation as 1st pellet and the supernatant as 1st Sup. Repeated treatments produced 2nd Sup and 2nd pellet. After 1000°C for 1 h the 2nd pellet produced solid as carbon. This sample washed with 4M HNO_3 , the supernatant as Acid Washing and the solid as Carbon Washed. The radioactivity assays clearly suggested that there had been 13% loss radioactivity due to the physical transfers (powder sticking on reactor tubes), only 1% loss was due to extensive acid treatment

(this standard acid digest with sonication would remove externally bound ^{99m}Tc , if any). As a result, the overall radioactivity retention is about 77% in the non-acid leachable solid sample. The surprisingly low activity in the acid wash solution suggests that majority of the particles were chemically protected with acid impenetrable graphitic jackets as indicated from the TEM micrographs.

10 ***Characterisation of the Product:***

High resolution TEM micrographs indicate that the iron-containing particles were mainly single nano-crystals of $\gamma\text{-Fe}$ metal (confirmed by XRD) although Fe_3C crystals as the inner cores were sometimes observed from the TEM lattice imaging. Detailed examination of the quasi-spherical carbon structures showed that in many cases, carbon lattice fringes (about 0.34 nm) could be traced, continuously yielding, quite surprisingly, hollow concentric carbon shelled structures. No prolonged exposure of the selected area to the electron beam was ensured (<60 seconds); hence the possibility of onions structure formation due to electron beam illumination is rejected. We found some complete filling but more frequently partial filling of the encapsulated iron particles to the carbon cages and there was no obvious preferred lattice fringe orientation with respect to the particle and the carbon layers. It is interesting to note that these highly ordered quasi-spherical concentric graphitic shells structures produced from iron cyanide decomposition are clearly different from tubular or polyhedral carbon structures. Formation of tubular/filamental graphitic carbons is well known to

occur when a carbon source (i.e. hydrocarbon gas) is in contact with iron particles at high temperatures [15]. Polyhedral nanoparticles of central cavities of varying sizes were also observed when amorphous carbons were exposed to high temperatures [16]. It is also noted that extremely small quantities of γ -Fe nano-metal crystallites encapsulated in polyhedral nanoparticles [17] were produced in a large amount of carbonaceous debris using the 'stuffed anode' modified Krätschmer-Huffman method where significant quantities of carbon nanotubes or 'sea urchin' structures [18] (carbon nanotubes grow radially from the metal nanocrystal) were also found. None of these structures were however seen in our sample. We note that our quasi-spherical graphitic shell structures rather closely resemble to the 'carbon onions' (see present Figure 10) reported by Ugarte who applied an intense electron beam irradiation on carbon nanotubes for the conversion [19], although structurally less perfect than his. The perfect onion is thought to compose of giant fullerenes giving perfect concentric shells with a quasi-spherical structure. Subsequent work [20] demonstrated 'carbon onions' are highly unstable when not being irradiated in an electron beam and will collapse into a disordered, quite often faceted, configuration, though still with a spheroidal structure somewhat akin to the structures observed.

Here, we have shown that heating the iron(oxo) cyanide nanoparticles (with or without foreign atoms) at 900°C results in graphitisation and the formation of many iron containing nanoparticles encapsulated in quasi-spherical graphitic shell. Presumably, the iron nanoparticles were formed from the reduction of iron

oxide with the nearby cyanide source during the carbonisation process.

The phenomenon of 'catalytic' graphitisation is a well-known one [11], but the mechanism in which graphitisation is promoted by the presence of a second phase is not well understood. In some cases, it is believed that the carbon is dissolved in the metal or metal carbide and re-precipitated as graphite. In other cases the metal or metal carbide particles may simply act as templates for the epitaxial growth of graphite. However, in all these cases, only long graphitic tubular or filamental forms of carbons and polyhedral nanoparticles are formed. The exclusive formation of highly ordered quasi-spherical graphitic nanocapsule structures indicative of 'fullerene-like structure' in our case however, has not been reported. It is believed that some nitrogen (residue of cyanide decomposition) resulting from the decomposition of cyanide species may play an important role in self-assembling the carbon atoms into corresponding quasi-spherical structure under our reaction conditions. Nitrogen (non-sp²) adopts different structure from carbon (sp²). Hence its incorporation into carbon graphene layer (all C is in sp² hybridization) is thought to be responsible for the curvature of the graphitic planes resulting in the particle encapsulation.

The present method for the formation of enclosed graphitic carbon structure embracing iron nanoparticle at relatively very mild temperatures (900°C) is unprecedented. Thus, iron filled spherical carbon nanocapsules of a very narrow size distribution in macroscopic quantities by the controlled decomposition of

cyanide-containing species could be prepared. Prior to pyrolysis, use of surfactant/polymer as stabiliser could provide a fine control of particle size. This is based on the observation that the polymer/surfactant adsorbs at the surface of the newly formed nano-particles through the formation of surface active micellar aggregates so preventing aggregation of nanoparticle in the solvent. The internal space is filled with an inorganic iron (oxo) cyanide nanoparticle aggregates with defined dimensions which are in turn controlled by the micellar dimensions. Further immobilisation of other foreign atoms onto the nanoparticle aggregates can be carried out. Here, we show that elemental Re element or Tc can be incorporated into the final product. Thus, the controlled carbonisation of these iron containing nano-assemblies provides the novel carrier which could open up a new avenue for preparing carbon nanocapsules filled with specific radioactive element(s).

There are considerable benefits of forming total carbon encapsulated radionuclides that are well suited for medical diagnosis and therapeutic purposes. The potential beneficiaries will be to patients with lung disease using the carbon encapsulated ^{99m}Tc -Fe nanoparticles described herein. For example, pulmonary embolism is a major cause of morbidity and mortality. It is estimated to account for approximately 21,000 deaths annually in England and Wales and over 200,000 in the USA. Early treatment by anticoagulation could save life, but it can sometimes be hazardous; a reliable means of diagnosis is therefore essential. The new technology for producing diagnostic or therapeutic material could provide rapid and efficient diagnosis suitable for

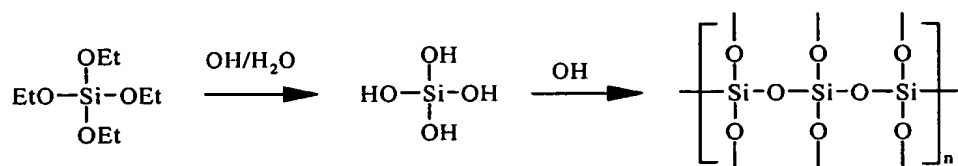
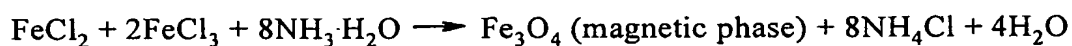
patients with this or other pulmonary disease, hence improving life expectancy and quality.

For Examples 5 and 6, iron (II) chloride tetrahydrate, iron (III) chloride hexahydrate, cobalt (II) chloride hexahydrate, tetraethyl orthosilicate (TEOS), cetyl trimethyl ammonium bromide (CTAB) were obtained from Aldrich in analytical grade quality and used without further purification.

Example 5: Preparation of the silica coated magnetic nanoparticle:

Formation of an aqueous microemulsion was carried out using de-ionized water, in organic solvent (i.e. dioctyl ether) and surfactant(s) (CTAB or oleic acid). As the size of such a micelle system is related to the ratio water/surfactant, a small amount of water to surfactant was used in order to reduce the size of the final nano-composite synthesized. The experiment was carried out in room temperature. The microemulsion was formed as follows: 6.0810g (0.01668 mol) CTAB was added into the 120ml dried organic solvent under vigorous stirring. After a well-distributed suspension of the surfactant in the solvent was achieved, 4.3ml aqueous solution containing 0.3428g $\text{FeCl}_2 \cdot 4\text{H}_2\text{O}$ and 0.9321g $\text{FeCl}_3 \cdot 6\text{H}_2\text{O}$ was added slowly in droplets into the suspension of the surfactant in solvent with nitrogen bubbled for two hours to avoid the possible oxidation of particles. After stopping of pumping of nitrogen, the system was stirred overnight to form the microemulsion. Then, nitrogen was bubbled again and 1.2ml 18.1M $\text{NH}_3 \cdot \text{H}_2\text{O}$ was placed in the path in which the N_2 gas flowed. One hour after the whole reaction system has turned black,

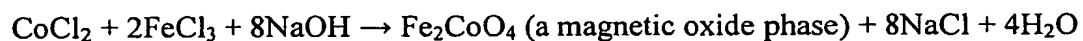
6.9351g TEOS was added into the reaction mixture. Nitrogen was continuing bubbled through this microemulsion for one more hour. Formation of the silica-gel coating can be achieved because the excess ammonia catalyzed hydrolysis/condensation of the organic soluble TEOS on the surface of the magnetic oxide nanoparticle (once in contact with water at the interface between the microemulsion and the bulk organic solvent). The envisaged chemical reactions are:



Ageing may take some time (i.e. 5 days) to form the silica layer fully. Then, when 30ml ethanol was added to the microemulsion system, precipitates of silica coated magnetic oxides were rapidly deposited by magnetic separation using an external magnetic field. The solvent was removed followed with an addition of another 30 ml ethanol into the precipitates and reflux the mixture for overnight to remove the surfactant from the precipitates. The precipitates were isolated again by magnetic separation and the solvent was removed. Ethanol, water and acetone wash of the precipitates several times. Finally the precipitates were dried in room temperature and a deep brown powder was collected for further characterization. These precipitates showed very strong magnetic force.

Example 6: Synthesis iron cobalt oxide core nanoparticle

The same microemulsion system as in Example 5 was applied to synthesize the iron cobalt oxide core nanoparticles. 6.0810g (0.01668 mol) CTAB was added into the 120ml dried organic solvent under vigorous stirring. After a well-distributed suspension of CTAB was achieved, 3.8ml aqueous solution contained 0.4056g $\text{CoCl}_2 \cdot 6\text{H}_2\text{O}$ and 0.9216g $\text{FeCl}_3 \cdot 6\text{H}_2\text{O}$ was added slowly in droplets into the suspension of CTAB in organic solvent with nitrogen bubbled two hours for avoiding possible oxidation of the particles. After stopping of pumping of the nitrogen, the system was stirred overnight to form the microemulsion. The system was slowly heated to 65°C and maintained at that temperature. 1.7 ml 10 M NaOH solution was added with nitrogen bubbled again. One hour after the whole reaction system turned black, it was cooled down, and 6.9351g TEOS was added into the reaction mixture. Nitrogen was bubbled through this microemulsion for one more hour. In this case:



Ageing also took about 5 days to form the silica layer fully. Then, the same method as above was used to wash and separate the precipitates and collect them for further characterization. These precipitates also showed very strong magnetic force.

Product characterization:

For characterizing the two kinds of nanoparticles of Examples 5 and 6, X-ray powder diffraction (XRD) analysis of the samples was performed with copper $\text{K}\alpha_1$ of 1.5405Å

radiation. Transmission electron microscopy (TEM) analysis was used to characterize the particle size, structure, morphology and composition through direct imaging, electron diffraction and elemental analysis of selected area (EDS). A Philip CM20 microscope operating at 200 kV equipped by an energy dispersive spectrometer (EDS) was used. Samples were gently ground, suspended in isopropanol and placed on carbon-coated copper grid after the evaporation of the solvent. Electron micrographs and EDS analysis of selected area were taken.

XRD spectrum (Figure 13) shows the nano-scale iron oxide of Example 5. From comparison with published XRD data the composition of the iron oxide could either be Fe_2O_3 or Fe_3O_4 or mixed phases of them since their lattice parameters would not allow further differentiation. It is noted that even a pure phase might still have undergone further reactions (air oxidation) when the particles were exposed in air.

The average diameter particle size D was determined by the Debye-Scherrer formula using the half maximum width β of X-ray diffraction lines ($D = (\kappa\lambda)/(\beta_{(hkl)}\cos\theta)$). In this equation, the D is the size of nanoparticle, κ is the Scherrer constant (0.9), λ is the X-ray wavelength (1.54056 in our experiments), and the $\beta_{(hkl)}$ is the full width half maximum (FWHM) of the reflection hkl measured in 2θ and is the corresponding Bragg angle. For the powder of Example 5, the average particle size is found to be about 17.33 nm.

For the iron cobalt oxide core nanoparticles of Example 6, the XRD spectrum (Figure 12) confirms the formation of nano-scale iron cobalt oxide. The result

was compared with the published XRD database, and showed that the iron oxide is similar to CoFe_2O_4 . According to the Debye-Scherrer calculation the average particle size is about 19.93 nm.

5 Figure 15 shows the EDS analysis of the iron oxide core nanoparticles of Example 5. From the analysis results taking the corrected AFZ into account, we obtain the atomic ratios:

Fe : Si : = 25.39 : 14.02 : 60.59.

10 The calculated formula is $\text{Fe}_3\text{O}_{3.84} \cdot 1.66 \text{SiO}_2$ confirming that the pre-dominant phase of the core is Fe_3O_4 with silica coat(s). In line with the XRD analysis results, it was found that the average composition of the particles was indeed between Fe_2O_3 and Fe_3O_4 phase (lying
15 closer to the Fe_3O_4 phase).

Figure 16 shows the EDS analysis of the iron cobalt oxide core nanoparticles of Example 6. From the analysis results, we obtain the atomic ratios:

Fe : Co : Si : O = 10.11 : 5.54 : 22.96 : 61.40

20 The calculated formula is $\text{CoFe}_{1.83}\text{O}_{2.79} \cdot 4.14 \text{SiO}_2$. The concentration of silica seems to be higher when compared with the EDS analysis of the iron oxide core nanoparticles of Example 5, but the calculated composition is more or less similar to the phase obtained
25 from the XRD analysis within experimental error.

Figure 17 shows the particles of Example 5, clearly revealing the metal oxide core surrounded by silica.

References:

1. D. Jones, *New Scientist*, Nov. 3, 1966.
2. J.R. Heath, S.C. O'Brien, Q. Zhang, Y. Liu, R.F. Curl, H.W. Kroto, F.K. Tittel, and R.E. Smalley, *JACS* 109, 359, 1985.1.
- 5 3. S.A. Majetich, J.O. Artman, M.E. McHenry, N.T. Nuhfer and S.W. Staley, *Phys Rev. B* 48, 16, 845, 1993.
4. S. Seraphin, D. Zhou and J. Jiao, *J. Appl. Phys.* 80, 1, 1996.
5. Y. Saito in *Carbon Nanotubes: Preparation and Properties*, ed. T.W. Ebbesen, CRC Press, 1996.
- 10 6. Y. Saito, *Carbon* 33, 979, 1995.
7. V.P. Dravid et al. *Nature* 374, 602, 1995.
8. J.R. Heath, S.C. O'Brien, Q. Zhang, Y. Liu, R.F. Curl, H.W. Kroto, F.K. Tittel and R.E. Smalley, *JACS* 109, 359, 1985.
9. D. Ugarte, *Chem. Phys. Lett.*, 209, 99, 1993.
- 15 10. S.C. Tsang, P.J.F. Harris, M.L.H. Green, *Nature*, 362, 520, 1993.
11. S.C. Tsang, Y.K. Chen, P.J.F. Harris and M.L.H. Freen, *Nature*, 372, 159, 1994.
12. S.C. Tsang, Z. Guo, Y.K. Chen, M.J.H. Green, H.A.O. Hill, T.W. Hambley and P.J. Sadler, *Angew. Chem. Int. Ed. Engl.* 36, 2197, 1997.
- 20 13. P.J.F. Harris and S.C. Tsang, *Chem. Phys. Lett.* 293, 53, 1998.
14. P.J.F. Harris and S.C. Tsang, *Carbon* 36, 1859, 1998.
15. R.T.K. Baker, P.S. Harris, R.B. Thomas, R.J. Waite, *J. of Catal.*, 30, 86, 1973.
16. P.J.F. Harris and S.C. Tsang, *Phil. Mag. A* 76, 667, 1997.
17. J.J. Host, M.H. Teng, B.R. Elliott, J.H. Hwang, T.O. Mason, D.L. Johnson, 25 V.P. Dravid, *J. Mater. Res.*, 12, 1268, 1997.
18. S. Subramoney, R. Ruoff, D.C. Lorents, R. Malhotra, *Nature*, 366, 637, 1993.
19. D. Ugarte, *Nature*, 359, 707, 1992.
20. G. Lulli, A. Parisani, G. Mattei, *Ultramicroscopy*, 60, 187, 1995.
21. 'Biomimetic Materials Chemistry' (ed. S. Mann), VCH Press, New York, 30 1996, chapter 1 and 4 and references therein.
22. W.M. Burch, P.J. Sullivan and C. McLaren. *Nucl. Med. Commun.*, 1986, 7, 865-871; W.M. Burch, P.J. Sullivan, F.E. Lomas, V.A. Evans, C.J. McLaren,

- R.N. Arnot. *J. Nucl. Med.*, 1986, 27, 842-846; T.J. Senden, K.H. Moock, J.F. Gerald, W.M. Burch, R.J. Browitt, C.D. Ling and G.A. Heath, *J. Nucl. Med.*, 1997, 38, 1327-133.
23. C.W. Ng, J. Ding, L. Wang, L.M. Gan, C.H. Quek, *J. Phys. Chem. A* 2000,
5 104, 8814-8822.
24. E. Boellaard, A.M. van der Kraan, J.W. Geus, *Preparation of Catalysis VI - Scientific Bases for the Preparation of Heterogeneous Catalysis*, G. Poncelet et al. (Editors), Elsevier, Amsterdam, 1995, p. 931.
25. *Chemistry and Physics of Carbon*, edited by P.A. Thrower, v1-25, Marcel
10 Dekker, New York, Inc. 1977.